Chemistry 115 Name Key

Dr. Cary Willard

Quiz 4a (20 points) September 28, 2010

All work must be shown to receive credit

1. (6 points) Give the proper name for each of the following compounds
	1. Na3P sodium phosphide
	2. Al2S3 aluminum sulfide
	3. CF4 carbon tetrafluoride
	4. NiO nickel(II) oxide
2. (6 points) Write the correct formula for each of the following compounds
	1. Potassium iodide KI
	2. Nitrogen triiodide NI3
	3. Zinc chloride ZnCl2
	4. Titanium(III) fluoride TiF3
3. (2 points) Draw the correct lewis electron dot structure for an atom of sulfur



1. (6 points) Circle the larger of the atoms/ions in each group below
	1. Br or Ca
	2. C or Sn
	3. Mg or Mg+2

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Quiz 4b (20 points) September 28, 2010

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1. (6 points) Give the proper name for each of the following compounds
	1. Li2S Lithium sulfide
	2. SO3  sulfur trioxide
	3. ZnF2 zinc fluoride
	4. CrBr3 chromium(III) bromide
2. (6 points) Write the correct formula for each of the following compounds
	1. Sodium nitride Na3N
	2. Calcium nitride Ca3N2
	3. Dibromine tetroxide Br2O4
	4. Tin(IV) bromide SnBr4
3. (2 points) Draw the correct lewis electron dot structure for an atom of nitrogen



1. (6 points) Circle the larger of the atoms/ions in each group below
	1. Na or Cl
	2. S or O
	3. N or N-3

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Quiz 4c (20 points) September 30, 2010

1. (6 points) Give the proper name for each of the following compounds
	1. AgF silver fluoride
	2. Ba3N2 barium nitride
	3. VBr2 vanadium(II) bromide
	4. PCl 5 phosphorous pentachloride
2. (6 points) Write the correct formula for each of the following compounds
	1. Zinc oxide ZnO
	2. Sulfur tribromide SBr3
	3. Aluminum sulfide Al2S3
	4. Ferric iodide FeI3
3. (2 points) Identify an atom which is isoelectronic with Se-2

 Kr or krypton

1. (2 points) What is meant by the term ionization energy.

The amount of **energy required** to remove and electron from a gaseous atom in its elemental state.

1. (4 points) Circle the atoms with the higher ionization energy in each group below
	1. Br or Ca
	2. C or Sn

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Quiz 4d (20 points) September 30, 2010

1. (6 points) Give the proper name for each of the following compounds
	1. CdF2 cadmium fluoride
	2. Mg3P2 magnesium phosphide
	3. TiCl3 titanium(III) chloride
	4. PBr3 phosphorous tribromide
2. (6 points) Write the correct formula for each of the following compounds
	1. Silver sulfide Ag2S
	2. Xenon hexafluoride XeF6
	3. Barium iodide BaI2
	4. Cuprous oxide Cu2O
3. (2 points) Identify an atom which is isoelectronic with Ca+2

Ar or argon

1. (2 points) What is meant by the term ionization energy.

The amount of **energy required** to remove and electron from a gaseous atom in its elemental state.

1. (4 points) Circle the atoms with the higher ionization energy in each group below
	1. Na or Cl
	2. S or O